

The background is a solid blue color. It is decorated with several 3D-rendered spheres in purple, yellow, and blue, some of which are connected by thin grey lines, resembling a molecular structure. In the top right and bottom left corners, there are white grid patterns that look like a staircase or a corner of a grid.

## Potions WESO 2026

- We want to know who is here! Please introduce yourself in the chat: Name. School.
- Please mute your microphone to reduce background noise.
- We will leave time for Q & A. Please submit your questions via the Google form that will be shared in the chat.
- This presentation, along with the questions and answers, will be posted on the WESO website.

The background is a solid dark blue. In the center, a white rectangular border contains the main text. To the left of this border is a molecular model with a yellow sphere, a pink sphere, and a blue sphere connected by grey rods. To the right is another molecular model with an orange sphere, a purple sphere, and a blue sphere connected by grey rods. In the top right corner, there is a white grid pattern. In the bottom left corner, there is a white and blue checkered pattern.

# POTIONS WESO 2026

Event Supervisor:  
Anastasia Yocum



# Goals for this event:



**Basic Chemistry  
Concepts**



**Laboratory  
Skills**



**Laboratory  
Safety**



**Teamwork**



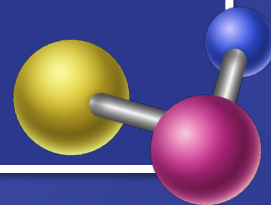
**Fun**



# EVENT DETAILS

- GRADES: 4 & 5
- TEAM SIZE: 1-2 participants (per grade)
- DURATION: 40 minutes (grades 4,5)

The detailed event descriptions and study guides can be found at:  
<https://wesoscience.org/events/>

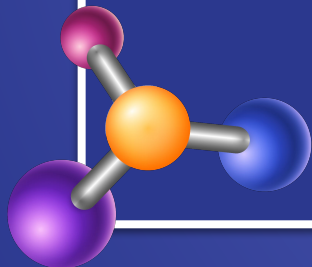




# EVENT FORMAT – Overview

## Two Components to this event :

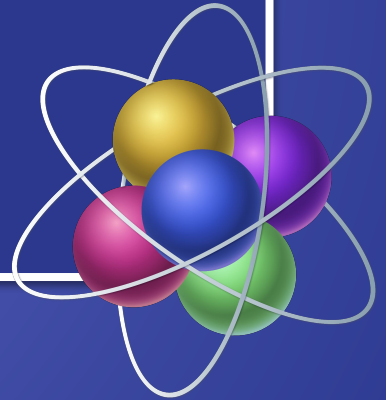
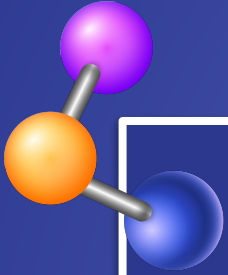
1. The test will mostly be experiment-based, where teams will conduct simple experiments, make observations and answer questions.
2. Relevant questions on basic chemistry concepts will be spread throughout the test.



REMEMBER - One test per team. They need to work together and collaborate on their answers!

# Event Format

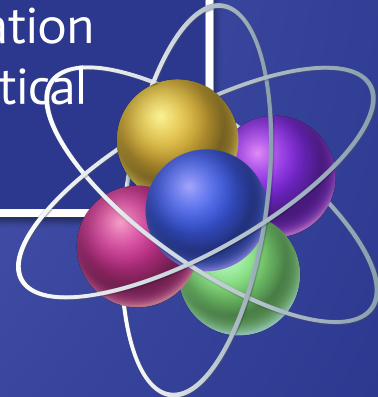
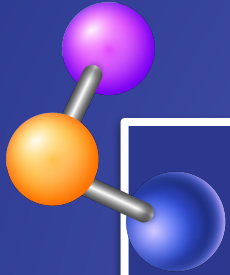
- Standard test section:
  - Multiple choice
  - Fill in the blank
  - True/False
  - Matching
  - Short answer

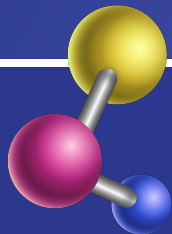


# Event Format

## Experimental section:

- Teams are given several small experiments, observations or tasks.
- All materials and detailed instructions will be provided.
- Teams will be asked questions (on paper) about the experiments and the concepts behind them.
- Most of this will be carried out at their team's work station but they may be asked to complete a portion at a practical skills area set up in the room.



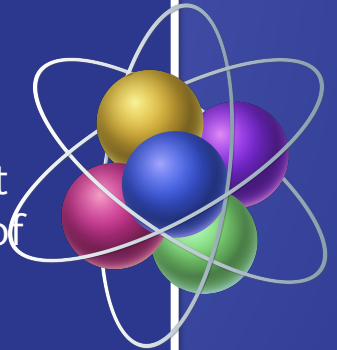


# Event Format

- Teams will be given one test packet which will have the questions for both sections in it. They are allowed to work in any order they wish, split it up (not recommended) or go back and forth.
- Answers will need to be placed on the answer sheet.
- Teams will be given both 10 and 5 minute warnings.
- Teams need to come to the event wearing proper lab safety equipment (long pants, closed toe shoes, lab coats or long-sleeved shirts over their WESO T-shirts, long hair pulled back and gloves and goggles worn when doing the experimental portion).
- Teams will be expected to adhere to lab safety rules and clean up their work area as instructed.

# SCORING

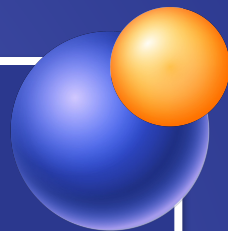
- Point value for each question will be clearly marked.
- Scoring will be based on the number of correct answers and how well the students follow lab safety and cleanup procedures.
- Completion time NOT considered in scoring.
- Teams will have points deducted if they do not practice proper lab safety rules and do not clean their work area properly.
- There will be several clearly marked tiebreaker questions at the end of the test which will only be graded in the event of a tie.



# Example of clean-up procedure:

- Continue to wear your gloves and goggles during clean-up.
- Leave all liquids in the rack on your workbench.
- Discard all weigh boats, spatulas, transfer pipets, pH strips and bags into the solid waste garbage (labeled).
- Periodic tables and pH scales should remain at your workstation.
- Remove your gloves and discard in the solid waste garbage.
- Leave goggles at your workstation if you did not bring your own.
- Raise your hand when you are finished. Your test will be collected, your workstation will be inspected, and you will be dismissed.

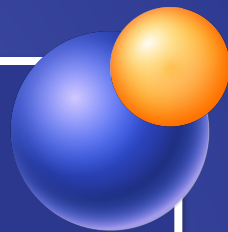
# EVENT CONCEPTS – Grades 4&5



- States of Matter: solid, liquid, gas
- Properties of solids, liquids, gases
- Transitions between phases: melting/freezing, evaporation/condensation, sublimation
- Law of Conservation of Mass
- Structure of matter: atoms, elements, molecules, compounds
- Basic atomic structure: Electrons, Protons, Neutrons, Nucleus
- Mixtures: Types of Mixtures (homogeneous, heterogeneous, solutions)

The study guide outlines the concepts and the level of comprehension expected.

# EVENT CONCEPTS – Grades 4&5



- Periodic table of elements: (Do not memorize! Students will be given a table)
  - Basic organization and information represented, understand the rationale behind periods and groups
  - General Understanding of characteristics of the families
  - Basic elements: H, O, N, C, Al, Na, Cu, Cl, P, S, He, Ar
    - General Characteristics of these elements
- Acids and Bases, pH scale
- Physical vs. Chemical Properties and Changes
- Ink Chromatography
- Principles and Practice of Filtration (solid and liquids)


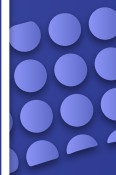
The study guide outlines the concepts and the level of comprehension expected.



# EVENT CONCEPTS – Grades 5 ONLY



## Additional topics for 5<sup>th</sup> Grade Only

- Chemical reactions
  - Ions (cations and anions) and Isotopes
  - Chemical Bonding (ionic vs. covalent)
  - Drawing a Bohr Model
  - Periodic Trends:
    - Atomic Radius Trends
    - Metallic Character Trends
  - Iodine Test
- 
- 

The study guide outlines the concepts and the level of comprehension expected.



# Periodic Tables

Periodic Table of the Elements																		5														
1 1A 11A H Hydrogen 1.008		2 2A 12A He Helium 4.003																13 3A 11B B Boron 10.811	14 4A 14A C Carbon 12.011	15 5A 15A N Nitrogen 14.007	16 6A 16A O Oxygen 15.999	17 7A 17A F Fluorine 18.998	18 8A 18A Ne Neon 20.180									
3 Li Lithium 6.941	4 Be Beryllium 9.012															5 Al Aluminum 26.982	6 Si Silicon 28.086	7 P Phosphorus 30.974	8 S Sulfur 32.06	9 Cl Chlorine 35.453	10 Ar Argon 39.948											
11 Na Sodium 22.990	12 Mg Magnesium 24.305	3 IB 3B Sc Scandium 44.956	4 IVB 4B Ti Titanium 47.88	5 VB 5B V Vanadium 50.942	6 VIB 6B Cr Chromium 51.996	7 VIIB 7B Mn Manganese 54.938	8 VIII 8 Fe Iron 55.833			9 VIII 9 Co Cobalt 58.933	10 VIII 10 Ni Nickel 58.693	11 IB 1B Cu Copper 63.546	12 IIB 2B Zn Zinc 65.38				13 Ga Gallium 69.723	14 Si Silicon 72.61	15 P Phosphorus 74.922	16 S Sulfur 78.96	17 Cl Chlorine 79.904	18 Kr Krypton 83.80										
19 K Potassium 39.098	20 Ca Calcium 40.078	21 Sc Scandium 44.956	22 Ti Titanium 47.88	23 V Vanadium 50.942	24 Cr Chromium 51.996	25 Mn Manganese 54.938	26 Fe Iron 55.833	27 Co Cobalt 58.933	28 Ni Nickel 58.693	29 Cu Copper 63.546	30 Zn Zinc 65.38	31 Ga Gallium 69.723	32 Ge Germanium 72.61	33 As Arsenic 74.922	34 Se Selenium 78.96	35 Br Bromine 79.904	36 Kr Krypton 83.80															
37 Rb Rubidium 85.468	38 Sr Strontium 87.62	39 Y Yttrium 88.906	40 Zr Zirconium 91.224	41 Nb Niobium 92.906	42 Mo Molybdenum 95.94	43 Tc Technetium 98.907	44 Ru Ruthenium 101.07	45 Rh Rhodium 102.906	46 Pd Palladium 106.42	47 Ag Silver 107.868	48 Cd Cadmium 112.411	49 In Indium 114.818	50 Sn Tin 118.71	51 Sb Antimony 121.76	52 Te Tellurium 127.6	53 I Iodine 126.905	54 Xe Xenon 131.29															
55 Cs Cesium 132.905	56 Ba Barium 137.327	57-71 Lanthanide Series		72 Hf Hafnium 178.49	73 Ta Tantalum 180.948	74 W Tungsten 183.85	75 Re Rhenium 186.207	76 Os Osmium 190.23	77 Ir Iridium 192.22	78 Pt Platinum 195.08	79 Au Gold 196.967	80 Hg Mercury 200.59	81 Tl Thallium 204.383	82 Pb Lead 207.2	83 Bi Bismuth 208.980	84 Po Polonium [209]	85 At Astatine [210]	86 Rn Radon [222]														
87 Fr Francium [223]	88 Ra Radium [226]	89-103 Actinide Series		104 Rf Rutherfordium [261]	105 Db Dubnium [262]	106 Sg Seaborgium [266]	107 Bh Bohrium [264]	108 Hs Hassium [277]	109 Mt Meitnerium [268]	110 Ds Darmstadtium [271]	111 Rg Roentgenium [272]	112 Cn Copernicium [285]	113 Nh Nihonium [284]	114 Fl Flerovium [289]	115 Uut Ununtrium [288]	116 Lv Livermorium [293]	117 Uus Ununseptium [294]	118 Uuo Ununoctium [294]														
Lanthanide Series																		57 La Lanthanum 138.905	58 Ce Cerium 140.12	59 Pr Praseodymium 140.908	60 Nd Neodymium 144.24	61 Pm Promethium 144.913	62 Sm Samarium 150.36	63 Eu Europium 151.965	64 Gd Gadolinium 157.25	65 Tb Terbium 158.905	66 Dy Dysprosium 162.50	67 Ho Holmium 164.930	68 Er Erbium 167.26	69 Tm Thulium 168.934	70 Yb Ytterbium 173.04	71 Lu Lutetium 174.967
Actinide Series																		89 Ac Actinium 227.028	90 Th Thorium 232.038	91 Pa Protactinium 231.036	92 U Uranium 238.029	93 Np Neptunium 237.048	94 Pu Plutonium 244.064	95 Am Americium 243.061	96 Cm Curium 247.070	97 Bk Berkelium 247.073	98 Cf Californium 251.083	99 Es Einsteinium 252.083	100 Fm Fermium 257.095	101 Md Mendelevium 258.1	102 No Nobelium 259.103	103 Lr Lawrencium [260]
Alkali Metal		Alkaline Earth		Transition Metal			Semimetal		Nonmetal		Basic Metal		Halogen		Noble Gas		Lanthanide		Actinide													

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<https://wesoscience.org/events/>

# STUDY GUIDE

- The study guide is posted on the WESO website under events.
- Provides information regarding the level of comprehension expected.
- Identifies differences in required topics between the grades.
- They contain sample questions.
- The study guide is meant for coaches not students!

# Basic Chemistry Lab Skills

- Use of the metric system: liquids, weight
- Identification and appropriate use of glassware and all components of the Potions Kit provided to the school
- Weigh out powders using electronic balance.
- Measure liquid volumes using graduated cylinders.
- Pour liquids into narrow-mouthed containers.
- Use a pipet/dropper to measure and transfer small amounts of liquid.
- Filter solutions using funnel and paper filters.
- Record observations of liquids and powders by examining their physical or chemical properties.



See the detailed event descriptions for more detail and the additional skills required for grades 4&5

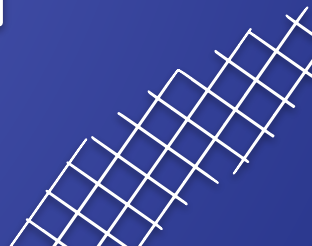
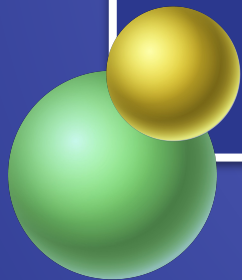
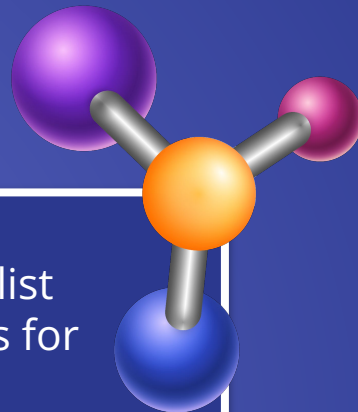
# SAFETY RULES

- Do not experiment on your own – always have an adult nearby.
- Read and follow all directions for your activity and use materials carefully.
- Read all warning labels on all materials being used.
- Wear protective clothing (lab coats, long pants, close toed shoes, gloves, and safety goggles/glasses).
- Long hair must be pulled back.
- When finished, clean up and dispose of all materials properly.
- Wash your hands after the activity.
- Never taste or directly smell any reagents.
- Keep materials away from your eyes.
- Never use any lab containers for food storage.
- Never eat or drink while conducting an experiment.
- Never play around/near chemicals.



# ADDITIONAL RESOURCES

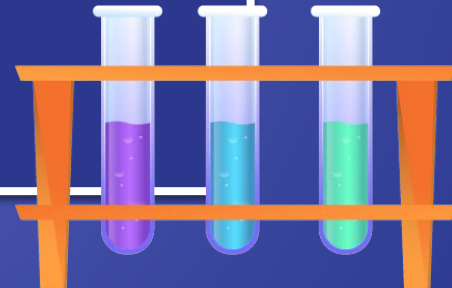
- The detailed event description and study guide list many websites that will serve as good resources for content and activities.
- There are many additional resources online.
- Make sure you screen websites etc. to make sure the information does not go beyond the required curriculum for this event.





## EVENT MATERIALS DISTRIBUTED BY WESO

- All schools have been provided with a Potions kit. Please contact your school's head coach in order to locate it.
- Full contents of the kit are listed in the event description.
- Some consumables are being replenished this year and will be given to your school's head coach.
- All necessary supplies will be provided by WESO for the competition day.
- The school is responsible for providing a balance for their coaching sessions.



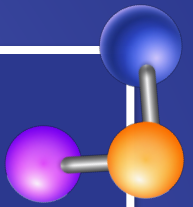






# GENERAL COACHING ADVICE

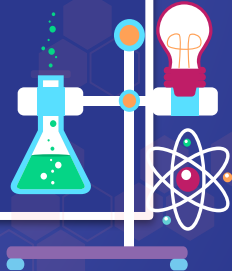
- Focus on the big picture.
- Keep it basic (We are not asking for the level of rigor required in a medical or research environment).
- Do not over coach your team.
- Do not answer questions with more information than requested.
- Focus on teamwork and problem solving.
- Make sure they support each other.
- Most importantly, make it fun!



# COACHING THE EXPERIMENTAL SECTION



-  Do as many hands-on activities that relate to the chemistry topics as possible.
-  Give them step-by-step instructions and make sure they read them thoroughly before they start.
-  Make sure your 2 students can work as a team and that they have a plan before they start.
-  They will not be given additional supplies or reagents if they make a mistake.



# SUGGESTIONS



## HELIUM BALLOON

- States of matter
- Elements (He)
- Volume



## MAKE KOOLAID

- Solutions
- Freeze it (states of matter and conservation of matter)

## CAN OF SODA

- Solutions
- Physical vs chemical change



## Clean Water Act

- Filter muddy water through sand





# Questions after tonight?

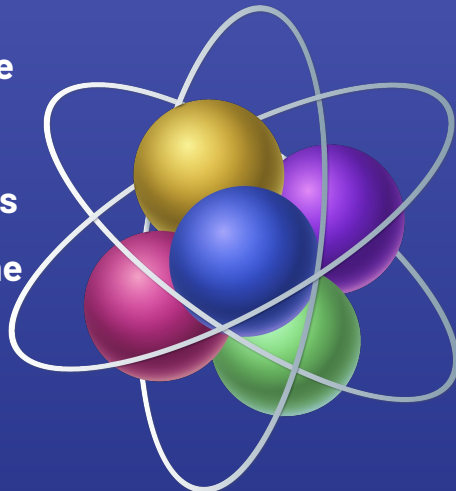
## New in 2026!

### WESO Discord Server

- Every WESO event will have its own channel
- Join the WESO server to submit your questions in the event chat
- Event supervisors or WESO board members will monitor the discussion and answer questions
- Event coaches can use the chat to exchange coaching ideas
- Go to [wesoscience.org/events/](https://wesoscience.org/events/) for details on how to join the WESO server and guidelines for its use

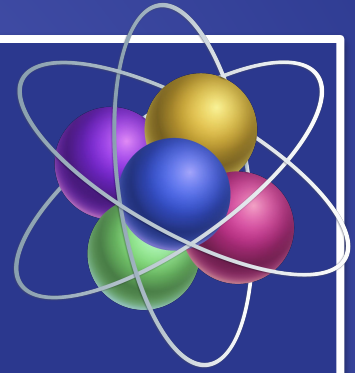


**Discord**





# QUESTIONS NOW?



- Please submit your questions now using the Google form that was shared with you in the chat.
- We will answer live and post all questions and written answers to the website following the meeting.

**Thank you for serving as an event coach and helping us bring back WESO to our community!**